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**WS 1.2 Density (SHOW ALL WORKS)**

1. A 13.2 mL coin weighs 4.8 g. Determine its density.  
Ans: \_\_\_\_\_

2. 138.42 g of salt water has a volume of 117.0 mL. Determine its density.  
Ans: \_\_\_\_\_

3. 0.46 g of hydrogen gas fills a 5.0 L bag. Determine hydrogen's density.  
Ans: \_\_\_\_\_

4. 25.2 mL of water are placed in a graduated cylinder. A 22.0 g cylinder is dropped in, and the water level rises to 32.4 mL. Determine cylinder density.  
Ans: \_\_\_\_\_

5. A graduated cylinder is placed on an electronic balance, and the scale reads 78.32 g. 10.0 mL of glycerine are added, and the scale reads 87.73 g. What is the density of glycerine?  
Ans: \_\_\_\_\_

6. A 3.0 cm x 4.0 cm x 6.7 cm die has a mass of 88.0 g. a) What is its density, and b) from what material is it most likely made?  
Ans: a) \_\_\_\_\_ b) \_\_\_\_\_

7. A cylinder has a mass of 528.6 g, a length of 14.2 cm, and a diameter of 2.30 cm. What is the cylinder most likely made of?  
Ans: \_\_\_\_\_

8. A bar has a mass of 120 g and a volume of 5.00 cm<sup>3</sup>. a) What is the bar's density? b) What is the bar's length in feet?  
Ans: a) \_\_\_\_\_ b) \_\_\_\_\_

9. a) How much would a 15.2 mL chunk of hydrogen weigh? b) How much would the same size chunk of osmium weigh?  
Ans: a) \_\_\_\_\_ b) \_\_\_\_\_

Answers are on bottom of next page.

Substance	Density
Aluminum	2.70 g/cm <sup>3</sup>
Brass	8.48 g/cm <sup>3</sup>
Carbon	2.26 g/cm <sup>3</sup>
Gold	19.30 g/cm <sup>3</sup>
Iron	7.87 g/cm <sup>3</sup>
Lead	11.34 g/cm <sup>3</sup>
Mercury	13.53 g/cm <sup>3</sup>
Platinum	21.45 g/cm <sup>3</sup>
Water	1.00 g/cm <sup>3</sup>
Wood	0.60 g/cm <sup>3</sup>
Zinc	7.14 g/cm <sup>3</sup>

**ELEMENTS AND THE PERIODIC TABLE** NAME \_\_\_\_\_

An element is a substance that \_\_\_\_\_ be \_\_\_\_\_ down into simpler substances by ordinary processes. Scientists know of \_\_\_\_\_ elements, only \_\_\_\_\_ of these are found in nature. The other elements are made in a \_\_\_\_\_, \_\_\_\_\_ elements are found in their \_\_\_\_\_ form. Most tend to \_\_\_\_\_ with other elements.

Oxygen (O) and Hydrogen (H) combine to form \_\_\_\_\_ (H<sub>2</sub>O).



Sodium (Na) and Chlorine combine to form \_\_\_\_\_ (sodium chloride - NaCl).

In the English language, we create thousands of words with \_\_\_\_\_ letters used in different combinations. In the same way, \_\_\_\_\_ natural elements combine to create a universe filled with billions of different things made of matter.

Sodium, hydrogen, carbon, and oxygen form common \_\_\_\_\_.

All atoms are made of \_\_\_\_\_, \_\_\_\_\_, and \_\_\_\_\_. It is the number of the \_\_\_\_\_ protons that distinguishes one element from another.

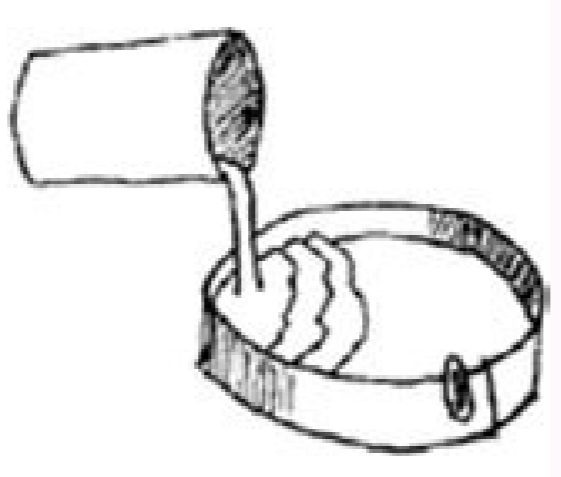
**States of matter**

- Easy Peasy**
1. Solids have a fixed shape. True or false?
  2. Which state of matter takes up all the space in a container?

- Medium**
3. Which state flows but does not fill it's container?
  4. Describe how the particles in a solid behave, using these keywords; flow, fill, fixed

- Hard**
5. Describe how the particles in a liquid behave.
  6. Put the states in order of the energy that the particles have.

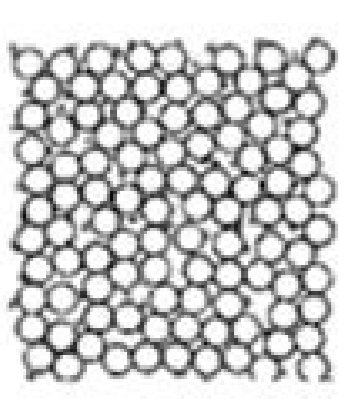
**Keywords**  
Solid, liquid, gas, particle, energy, flow, fill, container, fixed shape,



**Challenge Keywords**  
Kinetic energy  
Bonds  
Melting

- Very Hard**
7. Explain why solids have a fixed shape using the challenge keywords.
  8. Explain why we can store liquids in a bottle but not gases, using the challenge keywords.

- Nearly Impossible**
9. Explain why gases fill their container but liquids cannot (using the challenge keywords).
  10. If a material had stronger bonds, how would that affect the energy needed to change it from solid to liquid?









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